Molarity and Dilutions

Molarity = \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

How to make 500. ml of a 0.50 M methylene blue solution:

How would you make 250.0 ml of a 0.200 M solution of NaOH?

What is the molarity of a solution made by dissolving 3.00 g of NaCl in enough water to make 500.0 ml of solution?

How many moles of KBr are in 65.0 ml of a 0.300 M KBr solution?

M1V1 = \_\_\_\_\_\_\_ x\_\_\_\_\_\_\_\_

What is the concentration of a solution made by adding 50.0 ml of water to 20.0 ml of 0.500 M NaCl?

How would you make 750.0 ml of a 0.150M HCl solution starting with a stock solution of HCl that is 2.00 M?

Volume Percentage =

Mass Percentage =

PPM =

PPB =