Weak Acids and pKa

Notes

List of Strong Acids:

Definitions

 Strong Acids…

 Weak Acids…

HA + H2O 🡨🡪 A- + H3O+ Ka = pKa =

Write the reaction and equilibrium expression:

HNO2 + H2O

H2CO3 + H2O

What is the pH of a 0.200 M solution of formic acid? The Ka of acetic acid is 1.8 x 10-5.

A 0.0100 M solution of formic acid has a pH of 2.91. What is the Ka for formic acid? What is the pKa?

What is the Percent Ionization of a 0.500 M hypobromous acid (HBrO) solution? Ka = 2.3 x 10-9

A 0.050 M solution of a weak acid has a hydronium concentration of 6.5 x 10-4M.

What is the pH?

What is the percent ionization?

What is the Ka?