Weak Bases and Salts

Notes

Strong Bases:

Weak Bases:

Kb =

What is the pH of a 0.200M solution of ammonia? The Kb for ammonia is 1.76 x 10-5.

KaKb = Kw = pKa + pKb = pKa = pKb=

Benzoic acid (C6H5COOH) has a Ka of 6.25 x 10-5. What is the Kb and pKb for benzoate (C6H5COO-)?

What is the pH of a 0.050 M solution of sodium benzoate?

Salts

Cation, from strong base…

Cation, from weak base…

Anion, from strong acid…

Anion, from weak acid…

LiCO3 KBr NH4NO3

NH4HCOO