Nomenclature

Covalent compounds contain only

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Ionic compounds contain a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and/or a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.



Prefixes for Covalent Compounds

|  |  |
| --- | --- |
| Number | Prefix |
| 1 |  |
| 2 |  |
| 3 |  |
| 4 |  |
| 5 |  |
| 6 |  |
| 7 |  |
| 8 |  |
| 9 |  |
| 10 |  |

Names of Organic Hydrocarbons

|  |  |  |
| --- | --- | --- |
| Number of Carbons | Formula | Name |
| 1 |  |  |
| 2 |  |  |
| 3 |  |  |
| 4 |  |  |
| 5 |  |  |
| 6 |  |  |
| 7 |  |  |
| 8 |  |  |
| 9 |  |  |
| 10 |  |  |

Covalent Nomenclature

PCl5

N2O5

Ionic Nomenclature; cations of known charge

Groups I and II, \_\_\_\_\_\_\_\_\_\_\_\_, \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, and \_\_\_\_\_\_\_\_\_\_\_\_\_\_ always have the same charge.

NaCl

AlBr3

Mg3(PO4)2

Ionic Nomenclature; cations of variable charge

Fe2O3

Cr2(SO4)3

When using Latin names the lower oxidation state ends in \_\_\_\_\_\_\_ and the higher one ends in \_\_\_\_\_\_\_\_.

FeO

Hydrates have \_\_\_\_\_\_\_\_\_\_\_\_\_ molecules associated with the compund’s crystal structure.

CuSO4 •5H2O